CHEMICAL BONGING

½ Mark Questions

l.′	The Lewis dot s	tructure of Argon is				
	a) : Ar :	b) Ar:	c) Ar:	d) Ar.		
2.V	Vho proposed ic	onic bond?	3 1 3 7 2			
3.]	Formation of an	nion is favoured by		of an atom		
			ii) high electro	n affinity		
	iii) small ato	omic size	iv) low electro	negativy		
	a) only I	b) both I and i	i c) I, ii and iii	d) all of these		
1.F	or these format	tion of ionic bond the	electronegativity	difference between at	oms of two elements	
	should be					
	a) greater th	an or equal to 1.9	b) less t	han or equal to 0.9		
	c) equal to 1	5	d) equal	to 1		
5.V	Which elements	have octet configura	tion?			
5.V	Vhich one of the	e following wrong abo	out NaC <i>l</i> crystal?			
	a) It does no	t conduct electricity	in aqueous state			
	b) It is shou	ld be in water				
	c) The coord	ination number of c <i>l</i>	in NaC <i>l</i> is 6			
	d) it is face of	centered cubic crysta	1			
7.V	Which of the foll	owing statements is,	are not true			
	a) valence of	f metal is its group n	umber			
	b) Valence o	f non metal equal to	its group number			
	c) Noble gas are zero valent					
	d) Valence o	f non metal is 8- gro	up number			
3.	$S^{2-}: 2, 8, 8:: 1$					
).	Statement	I: Generally Noble g	gases have octet co	onfiguration		
	Statement	Statement II: helium gas duplet configuration				
	Which state:	ment is correct?				
١٥.	What is the ele	ctronic configuration	of Na+			
l 1.	Which type of b	oond will be formed v	vhen Magnesium r	eacts with Chlorine?		
12.	If an element X	belongs to II group a	and another eleme	nt Y belongs to VII gr	roup them what is th	
	formula of the	he compound?				
l3.	Which of the fo	llowing elements for	m unipositive in			
	a) Mg	b) Al	C) Ca	d) Na		
l4.	Which of the fo	llowing is not a coval	lent compound			
	a) BeCl ₂	b) BF ₃	c) Cacl ₂	d) CH ₄		
l 5.	The trivalent el	ement with electronic	c configuration 2,8	3,8 is		
	a) O ²⁻	b) Ca ²⁺	c) Mg ²⁺	d) P ³⁻		
16.	If the formula o	of metal chloride is $\it M$	Cl₃ then what is tl	ne formula of Metal S	Sulphate ?	

Assertion (A): Generally metals lose electrons to form positive ions				
Reason (R): Metals are electronegative in nature				
a) A and R are correct, R is correct explanation of A				
b) A and R are correct and R is not correct explanation of A				
c) A is correct and R is incorrect				
d) A is incorrect and R is correct				
. Match the following				
1) Sodium p) Mg^{2+}				
2) Magnisium q) S^{2-}				
3) Aluminium r) Na ⁺				
4) Sulphur s) Al ³⁺				
19. Which of the following has electron configuration 1s ² 2s ² 2p ⁶				
a) Na $^+$ b) F $^-$ c) Ne d) Mg $^{2+}$				
20.An element Na forms in ionic compound with another element Y.				
Then what is the change on the ion formed by X?				
21. What is formula of compound when A ³⁺ ion reacts with B ²⁻ ion?				
22. Chlorine : Valence 1 : : : valence 2				
i) Bromine ii) Mangisium iii) Sulphur iv) Calcium				
a) only I b) I and ii c) ii and iii d) ii, iii,iv				
23The ratio of coordination number of Na+ and Cl- ions in face centre cubic lattice crystal of NaCl				
a) 1:1 2) 6:1 c) 1:6 d) 3:4				
24. Match the following				
1) two metal atoms p) ionic bond				
2) Metal and non metal q) polar covalent bond				
3) Two similar non metals r)Metallic bond				
4) Two dissimilar non metals s) covalent bond				
ho proposed valance shell electron pair repulsion theory?				
26. The shape and bond angle in ammonia molecule is				
a) Pyramidal, 104 ⁰ 31' b) Tetrahydral, 107 ⁰ 48'				
c) Pyramidal, 107º48' d) Tetrahydral, 104º31				
What are the number of lone pairs and bond pairs present in water molecule?				
28.1 A ⁰ : 10 ⁻¹⁰ ::1 pm:				
29.Find odd one				
a) N ₂ – triple bond b) H ₂ -single bond				
c) O ₂ - double bond d) F ₂ -double bond				
The correct order of repulsions between electron pair is				
a) b.p-b.p> <i>l.</i> p- <i>l.</i> p>b.p- <i>l.</i> p				
c) b.p- <i>l.</i> p> <i>l.</i> p- <i>l.</i> p>b.p-b.p d) <i>l.</i> p- <i>l.</i> p>b.p-b.p-b.p-				
31. What is the number of covalent bonds present in methane?				
Becl2 : Linear : : : Tetrahydral				
a) BF ₃ b) CH ₄ c) H ₂ O d) NH ₃				
33. What are number of σ and π bonds present nitrogen molecule?				
34. What is the name of the molecule shown in the figure				

35. **Assertion (A):** Naphthalene is Soluble in kerosene Reason (R): Both Naphthalene and Kerosene are non polar in nature a) A and R correct and R is correct explanation of A b) A and R are correct and R is not correct explanation of A c) A is correct and R is correct d) A is incorrect and R is correct 36. What is the hybridization present in H₂O? 37. The number of lone pairs present in Sp3 hybrid orbital's in water molecule 3) 3 4) 6 1) 1 2) 2 38. Which overlap this figure represents? $(H\uparrow) + (H\downarrow) \longrightarrow (H\uparrow)H$ 39. What are the number of Hybrid orbital's formed in SP³ hybridization? 40.In Ammonia N-H bond; SP³-S:: in Becl₂, Be-Cl bond: _____ M.SRINIVASA RAO, SA(PS), AFC SCHOOL (AGKMHS), GUDIVADA. PH: 9848143855 Visit: srini science mind KEY 2. Kossel 3. C 4. A 1. a 5. Noble gases 7. B 6. a 8. 2,8 9. Both statements are correct 10. 2,8 11. Ionic bond 12. XY₂ 13. D 14) c 15) d 16) M2(SO4)3 17) c 18) 1-r, 2-p, 3-s, 4-q 19)I, ii, iii, iv 20) + 1 $21) A_2B_3$ 22) d 23) a 24) 1-r, 2-p, 3-s,4-q 25) sidwick, and powell 26) c 27) 2,2

28) 10⁻¹²m

33) 1σ , 2Π

38) S-S overlop

29) d

34) Ammonia

30) b

39) 4

31) 4

35) a

40) SP-P

32) b

36) SP³

37)2