

PERIODIC CLASSIFICATION

- Statement-I:** Dobereiner triad is based on atomic weight
Statement-II: F, Cl, Br form Dobereiner triad
Which of the following statements are true.
- Assertion (A) :** New elements are discovered due to Mendeleev's periodic table
Reason (R) : Mendeleev predicted that some elements were missing and left blank spaces at appropriate places in the table.
 - Both A and R are correct, R is the correct explanation of A
 - Both A and R are correct but R is not correct explanation of A
 - A is correct and R is incorrect
 - A is incorrect and R is correct.
- What is the name of element with atomic number 101?
- What is the atomic weight of divalent element with equivalent weight 4.5?
- The anomalous pair of elements in the following is
 - H and He
 - Ne and Ar
 - Kr and K
 - K and Ar
- Dobereiner triad : Atomic weight : : Modern periodic table : _____
- Match the following
 - Eka boron
 - Eka Aluminium
 - Eka Silicon
 - Gallium
 - Germanium
 - Scandium
- What are the number of periods and groups in modern periodic table !
- Find odd one
 - 1st period – shortest period
 - 7th period – incomplete
 - longest period – 4th period
 - 2nd period – 8 elements
- Electronic configuration of element in 2, 8, 7. Which of the following elements would it be chemically similar.
 - Nitrogen
 - fluorine
 - phosphorous
 - Argon
- What is the general electronic configuration of inert gases?
- Na : representative element ; _____ : Noble gas
 - He
 - Ne
 - Ar
 - Xe
 - Only I
 - Both I and ii
 - I, ii, iii
 - I, ii, iii, iv
- Which one of the following has electronic configuration $1s^2 2s^2 2p^6$?
 - Na⁺
 - Ne
 - F⁻
 - All of these
- To which group and period highest electronegative element belongs?
- 2nd period, 13rd group element : Boron : : _____ : Phosphorus
 - 3rd period, 14th group
 - 3rd period, 13rd group
 - 2nd period, 15th group
 - 3rd period, 15th group
- Which one of the following decreases in a group from top to bottom
 - atomic size
 - metallic nature
 - electro positivity
 - electro negativity
- 1 pm = _____ m
- Which one of the following has greatest size
 - Na
 - Na⁺
 - Mg²⁺
 - Al³⁺
- Find the wrong statement
 - Ionization energy decreases with increase in atomic size
 - Ionization energy increase with nuclear charge
 - As screening effect increases ionization energy increases
 - The unit of ionization energy is Kcal.Mol⁻¹

20. Which of the following is not a non metal
 a) Cl b) O c) H d) As
21. **Assertion(A):** First period is the shortest period.
Reason(R): It has only 3 elements
 a) A and R are correct, R is a correct explanation of A
 b) A and R are correct, R is not correct explanation of A
 c) A is correct but R is incorrect
 d) A is incorrect and R is correct
22. Which one of the following increases in a group from top to bottom
 A) atomic size b) metallic nature
 c) electropositivity d) electronegativity
23. Which one of the following elements is electronegative
 a) Na b) O c) Mg d) Ca
24. Be : Valence 2 : Chlorine : _____
 a) Valence 1 b) Valence 2 c) Valence 3 d) Valence 4
25. Which among these have smaller size?
 a) Na b) Na⁺ c) Mg²⁺ d) Al³⁺
26. Give two examples for metalloids?
27. The correct ascending order of atomic size of the following elements C, Li, N, Be
 a) C, Li, N, Be b) Li, Be, C, N c) N, C, Be, Li d) Be, Li, N, C
28. The correct order of electronegativity in the following elements is
 a) F > Cl > O b) F > O > Cl c) O > F > Cl d) Cl > F > O
29. Find odd one
 a) F b) Ca c) Br d) I
30. Which of the following statement is wrong.
 a) As screening effect increases ionization energy decreases
 b) As atomic size increases ionization energy decreases
 c) Ionization energy is expressed in KJ mol⁻¹
 d) Nitrogen has less ionization energy compared to oxygen.
31. 1 nm = _____ m
32. **Statement-I:** 4f elements are called lanthanides
Statement-II: s, p block elements except noble gases are called representative elements
 Which of the following statements is correct?
33. Modern periodic table is based on _____
 a) Atomic number b) Atomic size
 c) Electronic configuration d) Both a and c
34. Which inert gas does not have octet configuration?
35. d block elements : Transition elements : : _____ : Inner transition elements.
36. Find odd one based on group they belongs.
 a) F b) Cl c) O d) I
37. Which group elements are called halogens?
38. Matching
- | Group | Valence |
|---------------------------|----------------|
| 1) Valence of group I A | p) 4 |
| 2) Valence of group IV A | q) 3 |
| 3) Valence of group VI A | r) 1 |
| 4) Valence of group III A | s) 2 |

39. Where do Na and N belong?

- a) S-Block b) Na belongs to s-block and N belongs to p-block
c) Na belongs to p block and N belongs to s block d) p block

40. Match the following

- 1) Alkali Metals p) VI A or 16
2) Alkaline earth metals q) IA or I
3) Chalcogens r) VII or 17
4) Halogens s) II A or 2

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KEY

- 1) Statement I is correct and statement II is incorrect 2) a
3) Mandeleevium
4) 9 5) d 6) Atomic number 7) 1-r, 2-p, 3-q, 8) 7,18
9) C 10) b 11) $ns^2 np^6$ 12) d 13) d
14) 2nd period, 17th group 15) d 16) d 17) 10^{-12} m 18) a
19) C 20) D 21) C 22) a,b,c 23) b 24) a
25) d 26) Si,Ge 27) c 28) b 29) b 30) d
31) 10^{-9} m 32) Both statements are correct 33) d 34) He
35) f block element 36) c 37) VII A or 17th group
38) 1-r, 2-p, 3-s, 4-q 39) b 40) 1-q, 2-s, 3-p, 4-r

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